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Thermodynamic study of the gaseous molecules Al_2N , AlN , and Al_2N_2 by Knudsen cell mass spectrometry

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The Knudsen effusion mass spectrometric method has been employed to measure the equilibrium partial pressures of the Al_2N molecule over the AlN – Au –graphite system. Theoretical computations were carried out to determine the structure, molecular parameters, and thermodynamic properties of Al_2N . The partial pressures have been combined with the calculated thermal functions to determine the atomization enthalpy, $\Delta_a H_0^\circ$, and enthalpy of formation, $\Delta_f H_{298.15}^\circ$, in kJ mol^{-1} , of 783.2 ± 15 and 342.7 ± 15 for Al_2N , respectively. Upper values for the dissociation energy of AlN , $D_0^\circ(\text{AlN}, g) \leq 368 \pm 15 \text{ kJ mol}^{-1}$, and for the atomization enthalpy of Al_2N_2 , $\Delta_a H_0^\circ(\text{Al}_2\text{N}_2, g) \leq 1402 \text{ kJ mol}^{-1}$ have been obtained. These results are discussed and compared with recent theoretical literature values. © 2000 American Institute of Physics. [S0021-9606(00)01448-3]

INTRODUCTION

Aluminum–nitrogen clusters have received increased attention in the last few years because of their potential use as precursors in the growth of new materials with uncommon properties,¹ and because of gaining a better understanding of the fundamental mechanisms leading to the thin-film synthesis in order to optimize the deposition process.²

Aluminum nitride (AlN) is a useful material for substrates in microelectronics devices, due to its high thermal conductivity, very low electrical conductivity, moderately low dielectric constant, and a thermal expansion coefficient which is close to that of silicon.^{3,4} A recent study by Bhatti *et al.*⁵ showed that the use of AlN thin films as piezoelectric coating enables higher optical phase response over a wider frequency range as compared to similar devices using ZnO .

There are few experimental studies on gaseous aluminum nitrides. Gingerich⁶ observed for the first time the molecule Al_2N by Knudsen effusion mass spectrometry, and determined its atomization enthalpy, $\Delta_a H_{298.15}^\circ$, as $828 \pm 21 \text{ kJ mol}^{-1}$. He suggested the AlNAl arrangement of the atoms from the comparison of the enthalpies of atomization of Al_2N with those of Si_2N and Al_2O . Simmons and McDonald⁷ observed the emission spectrum of AlN in the visible region, and obtained the vibrational frequency of the AlN ground state as 747 cm^{-1} . Andrews *et al.*⁸ used infrared matrix isolation spectroscopy to study the products of aluminum atoms reacting with dinitrogen. They identified AlN_2 , Al_2N , Al_2N_2 , AlN_3 , and Al_3N molecules by nitrogen isotopic substitution and comparison with computed isotopic vibrational frequencies.

Several computational studies^{9–16} have been carried out on aluminum–nitrogen clusters to optimize their structures, to calculate the relative stability of different isomers, the vibrational frequencies and ionization energies, and to estimate their atomization energies.

In continuation of our recent investigations on small nitrogen containing group 13 and 14 clusters,^{17–21} we employed the Knudsen cell mass spectrometric method to de-

termine the atomization energies and enthalpies of formation of the Al_2N , AlN , and Al_2N_2 molecules. We also performed theoretical computations of Al_2N .

THEORETICAL INVESTIGATIONS

Ab initio calculations were carried out utilizing the GAUSSIAN 98 program package²² in order to gain information about the molecular parameters and physical–chemical properties of the Al_2N ground state.

The calculations were performed at the density functional level of theory using the Becke three-parameter exchange functional with the Lee, Yang, and Parr correlational functional (B3LYP). The basis set employed for this study was a correlated consistent polarized valence triple-zeta basis set, cc-pVTZ.

Ab initio calculations were performed on the three possible isomers of Al_2N , two linear structures, Al-N-Al and Al-Al-N , and a cyclic one. The results are listed in Table I. From the calculations the most stable isomer is found to have the linear symmetric structure with a $^2\Sigma_u^+$ electronic ground state. This result is also in agreement with experimental and theoretical results for similar molecules, Si_2N ,^{23,24} and B_2N ,^{25–28} which too have linear symmetric structures as the most stable ground state geometry. In Fig. 1 the potential energy of the ground state of Al_2N is plotted as a function of the AlN-Al bond distance at the B3LYP/cc-pVTZ level of theory.

The asymmetric linear structure, Al-Al-N , with a $^2\Sigma^+$ electronic ground state lies about 375 kJ mol^{-1} ($31\,350 \text{ cm}^{-1}$) above the symmetric Al_2N ground state, and the cyclic structure with the same electronic multiplicity as that of Al-N-Al is optimized as the linear symmetric one. The lowest energy cyclic structure has a 4B_2 electronic state and lies about 237 kJ mol^{-1} ($19\,810 \text{ cm}^{-1}$) above the ground state of the linear symmetric Al_2N .

The vibrational frequencies for the three vibrational normal modes, the optimized molecular parameters, and the atomization enthalpies, $\Delta_a H_0^\circ$, of the $X^2\Sigma_u^+$ state of the linear

TABLE I. The optimized bond distances, vibrational frequencies, electron affinity (EA), ionization energy (IE), and atomization enthalpy ($\Delta_a H_0^0$) of the ${}^2\Sigma_u^+$ ground state of the linear symmetric Al₂N and of its isomers at the B3LYP/cc-pVTZ level of theory.

Electronic state	Symmetry	Total energy ^a	$r_{\text{Al-N}}$ ^b	$r_{\text{Al-Al}}$	α_{AlNAl}	ν_1	ν_2	ν_3	EA ^c	IE	$\Delta_a H_0^0$ ^d
${}^2\Sigma_u^+$	$D_{\infty h}(\text{AlNAl})$	-539.663 409 2	1.726		180	527 ^e	94	1061	2.27	8.01	731.4
${}^2\Sigma^+$	$C_{\infty v}(\text{AlAIN})$	-539.520 527 6	1.692	2.723	180	251	29	872			359.8
4B_2	$C_{2v}(\text{Al}_2\text{N})$	-539.573 320 2	1.817		91.5	690 ^f	307	706			495.3

^aThe computed total energy is in hartree.

^bThe bond length is in Å.

^cEA and IE are in eV.

^d $\Delta_a H_0^0$ is in kJ mol⁻¹.

^e ν_1 is the symmetric stretching (σ_g), ν_3 the antisymmetric stretching (σ_u), and ν_2 the bending vibrational mode (π_u), in cm⁻¹.

^f ν_1 is the symmetric stretching (a_1), ν_3 the antisymmetric stretching (b_1), and ν_2 the bending vibrational mode (a_1), in cm⁻¹.

symmetric isomer and of the lowest lying electronic states of the other isomers of Al₂N are listed in Table I. For the most stable isomer the electron affinity (EA) and the ionization energy (IE) are also given. The computed EA, IE, and $\Delta_a H_0^0$ reported in Table I were corrected for the zero point energies (ZPE). Our computed values for the symmetric (σ_g) and antisymmetric (σ_u) stretching vibrational modes agree well with the corresponding experimental values of 957 cm⁻¹ and 545 cm⁻¹ by Andrews *et al.*⁸ The theoretical calculations by Andrews *et al.*⁸ at the B3LYP level of theory using 6-311 + G* and cc-pVDZ basis sets give results consistent with ours.

In order to support our computations, we also calculated the bond distances, vibrational frequencies, electron affinity, ionization energy, and atomization enthalpy ($\Delta_a H_0^0$) for the AINC molecule at the same level of theory. The results of these calculations show good agreement with the experimental values of bond distances,²⁹ vibrational frequencies,³⁰ and atomization energy.²⁰

EXPERIMENTAL PROCEDURE

The mass spectrometer and experimental procedures used for this work have been described previously.³¹ The sample of finely powdered AlN, graphite, and gold was vaporized from a high density graphite Knudsen cell. The measurements were performed following the investigation of the AINC/AICN molecule,²⁰ and divided into series 1 and series 2, because of the different ionization energies used.

The cell was heated by radiation by a tungsten coil resistor and the temperatures were measured with a calibrated Leeds and Northrup optical pyrometer focused onto a black body hole at the bottom of the cell. The ions were produced with ionizing electrons and electron emission current of 30 eV and 1 mA for series 1, and 15.5 eV and 1 mA for series 2. The acceleration voltage used was 3 kV; the electron multiplier was operated at 1.5 kV for both series.

The ions pertinent to the present investigation were Al⁺, Au⁺, AlAu⁺, AINC/AICN⁺, Al₂N⁺, and AlN⁺. The ions

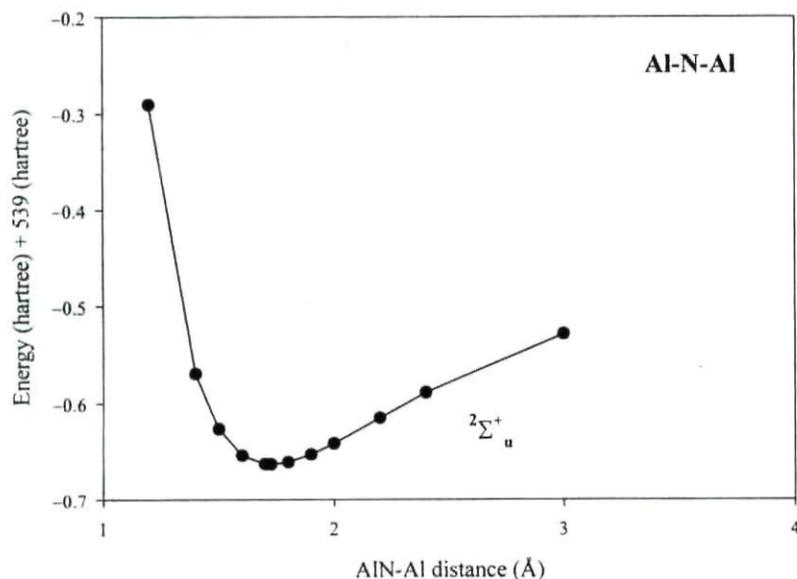


FIG. 1. Potential energy as a function of the AlN-Al distance of the ${}^2\Sigma_u^+$ ground state of Al₂N, calculated at the B3LYP/cc-pVTZ level of theory.

TABLE II. Measured ion currents, in A, and third-law values, in kJ mol^{-1} , of the enthalpy of reactions (1) and (2), $\Delta_r H_0^\circ$, and enthalpy of atomization, $\Delta_a H_0^\circ$, of Al_2N .

T (K)	Ion intensities					$\Delta_r H_0^\circ$ reaction 1	$\Delta_a H_0^\circ$	$\Delta_r H_0^\circ$ reaction 2	$\Delta_a H_0^\circ$
	Al_2N^+	Al^+	AuAl^+	Au^+	AlNC^+				
Series 1									
1818	$9.22E-13$	$1.13E-7$	$2.45E-9$	$8.70E-10$	$1.25E-9$	-35.6	803.5	319.2	793.2
1862	$1.50E-12$	$1.99E-7$	$3.40E-9$	$1.20E-9$	$2.89E-9$	-42.0	797.1	327.9	784.5
1862	$8.53E-13$	$2.02E-7$	$3.40E-9$	$1.20E-9$	$2.80E-9$	-50.3	788.8	336.9	775.5
1857	$7.29E-13$	$2.08E-7$	$4.44E-9$	$1.43E-9$	$2.99E-9$	-54.9	784.2	338.9	773.5
1805	$3.49E-13$	$9.74E-8$	$3.15E-9$	$1.10E-9$	$9.63E-10$	-46.2	792.9	329.3	783.1
1940	$1.20E-11$	$7.16E-7$	$1.13E-8$	$2.75E-9$	$1.33E-8$	-40.6	798.5	328.4	784.0
						-44.9 ± 7.0^a	794.2 ± 7.0	330.1 ± 7.1	782.3 ± 7.1
Series 2									
1886	$1.24E-12$	$3.12E-07$	$5.70E-09$	$1.03E-09$	$3.19E-09$	-53.4	786.5	339.8	772.6
1907	$1.58E-12$	$4.35E-07$	$6.70E-09$	$1.18E-09$	$4.79E-09$	-57.0	782.8	345.0	767.4
1922	$2.05E-12$	$5.61E-07$	$7.92E-09$	$1.36E-09$	$6.43E-09$	-58.3	781.5	347.5	764.9
1940	$3.41E-12$	$6.79E-07$	$8.60E-09$	$1.43E-09$	$9.00E-09$	-56.6	783.2	345.5	766.9
						-56.3 ± 2.1	783.5 ± 2.1	344.5 ± 3.3	768.0 ± 3.3

^aThe error terms are standard deviations.

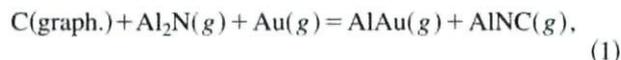
were identified by their mass-to-charge ratio, isotopic abundance, and, where possible, by their ionization efficiency curves. At each measurement a movable slit was interposed into the molecular beam to distinguish between ions produced from species in the beam and from residual gases with the same mass-to-charge ratio in the ionization region of the mass spectrometer. The electron impact energy scale was calibrated against the first ionization potential of Au (9.22 eV) (Ref. 32) and the appearance potential of (8.1 ± 0.5) eV was measured for Al_2N by the extrapolated voltage differences method.³³ The measured appearance potential agrees with our computed value of 8.01 eV (see Table I), and with the generalized gradient approximation (GGA) value of 7.84 eV.¹¹ The ion current data of the species pertinent to this investigation are listed in Table II.

The ion intensities I_i measured for the various species were converted into partial pressures by the relation $p_i = k_i I_i T$, where $k_i = k/(\sigma n \gamma)$ is the pressure calibration constant, k is the instrumental constant, σ is the ionization cross section, n is the isotopic abundance, and γ is the multiplier gain. The instrumental constant was determined from the known $\text{AlAu}(g) = \text{Al}(g) + \text{Au}(g)$ equilibrium, using the measured corresponding ion currents and the known dissociation energy, $D_0^\circ(\text{AlAu}, g) = (322.2 \pm 6.3) \text{ kJ mol}^{-1}$.³⁴ The resulting instrumental constant, in $\text{bar A}^{-1} \text{K}^{-1}$, is 26.6 for series 1, and 24.3 for series 2. The uncertainty of k is estimated to be about 20%.

RESULTS AND DISCUSSION

Al_2N

For the determination of the enthalpy of formation, $\Delta_f H_0^\circ$, and the enthalpy of atomization, $\Delta_a H_0^\circ$, of Al_2N , the enthalpy changes of the following pressure independent reactions,



were evaluated according to the third-law method, using the relation $\Delta_r H_0^\circ = -RT \ln K_p - T\Delta[(G_T^\circ - H_0^\circ)/T]$.

The Gibbs energy functions, $(G_T^\circ - H_0^\circ)/T$ (GEF_0), and the heat content functions, $H_T^\circ - H_0^\circ$ (HCF_0), needed in the evaluation of the reactions enthalpies were taken from literature for $\text{Al}(g)$,³⁵ $\text{AlN}(s)$,³⁵ $\text{C}(\text{graph.})$,³⁶ $\text{Au}(g)$,³⁷ AlAu ,³⁴ and AlNC .²⁰ Those for $\text{Al}_2\text{N}(g)$ were computed according to statistical thermodynamic procedures, using the rigid-rotator harmonic-oscillator approximation³⁸ and the molecular constants derived in the present theoretical investigation, bond distance, $r_{\text{Al-N}} = 1.726 \text{ \AA}$, and bending vibrational mode, $\nu_2 = 94 \text{ cm}^{-1}$, for which there are no experimental values. For the stretching vibrational modes the experimental values of $\nu_1 = 545 \text{ cm}^{-1}$ and $\nu_3 = 957 \text{ cm}^{-1}$ by Andrews *et al.*⁸ were used. In analogy to B_2N ,²⁷ the $2\Sigma_g^+$ excited state with an estimated transition energy (T_e) value of 6000 cm^{-1} was also considered. Table III lists the thermal functions of Al_2N .

Reactions (1) and (2) are pressure independent, therefore the measured ion intensities listed in Table III were utilized directly to calculate the equilibrium constant of the reactions (1) and (2), after correcting them for the isotopic abundance, multiplier response and ionization cross sections as described

TABLE III. The Gibbs energy functions, $(G_T^\circ - H_0^\circ)/T$ (GEF_0), in $\text{JK}^{-1} \text{mol}^{-1}$, and the heat content functions, $H_T^\circ - H_0^\circ$ (HCF_0), in kJ mol^{-1} , for the Al_2N , AlN , and Al_2N_2 molecules.

Species	Temperature (K)						
	298.15	1400	1600	1800	2000	2200	
Al_2N	-GEF ₀	221.0	299.2	306.7	313.5	319.6	325.2
	HCF ₀	13.21	78.57	91.08	103.7	116.4	129.3
AlN	-GEF ₀	198.9	250.7	255.4	259.6	263.4	266.9
	HCF ₀	9.264	49.68	57.18	64.70	72.23	79.78
Al_2N_2	-GEF ₀	239.5	334.2	344.0	352.9	361.0	368.5
	HCF ₀	14.72	101.6	119.3	137.2	155.3	173.4

in Ref. 20. The multiplier gain, γ , for Al₂N was measured with a 50% transmission grid as 10⁵ times 3.87. The relative ionization cross sections, σ , in 10⁻¹⁶ cm², used were at 30 eV: Al₂N, 15.2 as 0.75 \times (2 $\sigma_{\text{Al}} + \sigma_{\text{N}}$), where σ_{Al} and σ_{N} were taken from Freund *et al.*³⁹ and Brook *et al.*,⁴⁰ respectively; and at 15.5 eV: Al, 9.27;³⁹ Au, 3.10;⁴¹ C, 0.70;⁴⁰ N, 0.25;⁴⁰ Al₂N, 14.1 as 0.75 \times (2 $\sigma_{\text{Al}} + \sigma_{\text{N}}$); AlAu, 9.28 as 0.75 \times ($\sigma_{\text{Al}} + \sigma_{\text{Au}}$); AlNC, 7.66 as 0.75 \times ($\sigma_{\text{Al}} + \sigma_{\text{N}} + \sigma_{\text{C}}$).

The literature values for the enthalpies of formation, $\Delta_f H_0^\circ$, in kJ mol⁻¹, used to obtain the final values for the enthalpy of formation and atomization of Al₂N were: 368.0 \pm 1.3,³⁷ 711.2 \pm 0.5,³⁶ 470.8 \pm 0.1,³⁶ 327.6 \pm 3.0,³⁵ -341.3 \pm 3.0,³⁵ 372.9 \pm 8.4,³⁴ 281.3 \pm 14,²⁰ for Au(g), C(g), N(g), Al(g), AlN(s), AlAu(g), and AlNC(g), respectively. From the $\Delta_f H_0^\circ$ values, see Table III, and these auxiliary literature data the respective $\Delta_f H_0^\circ$ and $\Delta_a H_0^\circ$ were derived, using the following relations: for reaction (1), $\Delta_f H_0^\circ(\text{Al}_2\text{N}, g) = \Delta_f H_0^\circ(\text{AlAu}, g) + \Delta_f H_0^\circ(\text{AlNC}, g) - \Delta_f H_0^\circ(\text{Au}, g) - \Delta_f H_0^\circ$ and $\Delta_a H_0^\circ(\text{Al}_2\text{N}, g) = 2\Delta_f H_0^\circ(\text{Al}, g) + \Delta_f H_0^\circ(\text{N}, g) - \Delta_f H_0^\circ(\text{Al}_2\text{N}, g)$; for reaction (2), $\Delta_f H_0^\circ(\text{Al}_2\text{N}, g) = \Delta_f H_0^\circ + \Delta_f H_0^\circ(\text{Al}, g) + \Delta_f H_0^\circ(\text{AlN}, s)$.

By giving the same weight to both reactions employed in the present investigation we obtain the values, in kJ mol⁻¹: 783.2 \pm 15, 790.0 \pm 15, 342.8 \pm 15, and 342.7 \pm 15 for $\Delta_a H_0^\circ$, $\Delta_a H_{298.15}^\circ$, $\Delta_f H_0^\circ$, and $\Delta_f H_{298.15}^\circ$ of Al₂N(g), respectively. Here the uncertainties are the overall errors obtained as discussed in Schmude *et al.*⁴² Our computed $\Delta_a H_0^\circ(\text{Al}_2\text{N}, g)$ of 731 kJ mol⁻¹ or 7.58 eV is about 7% smaller than the experimental value of 783.2 kJ mol⁻¹; the value of 758 kJ mol⁻¹ or 7.86 eV from Nayak *et al.*,¹¹ obtained using the generalized gradient approximation (GGA) level of theory, is about 3% smaller than our experimental value. The value of 728 kJ mol⁻¹ or 7.55 eV by Andrews *et al.*⁸ based on B3LYP/cc-pVDZ calculations without ZPE correction practically coincides with our computed value.

The early $\Delta_a H_0^\circ(\text{Al}_2\text{N}, g)$ of 822 \pm 21 kJ mol⁻¹ (Ref. 6) differs from our result because of the different assumptions in the thermal functions used and of the different partial pressures obtained in the two investigations. A representative example of this difference is given for reaction (1) at 1940 K. At this temperature the difference of the Δ_{GEF_0} used in this study and those in Ref. 6 is 5.9 J mol⁻¹ K⁻¹, and $\ln K_p$ obtained in this investigation is 0.8 higher than $\ln K_p$ in Ref. 6. These differences reflect on an increase of the enthalpy change of reaction (1) for Ref. 6 of about 24 kJ mol⁻¹, and thus the corresponding atomization energy is lowered by the same amount. The corrected value of 798 \pm 21 kJ mol⁻¹ is within the error limits of the present investigation.

AlN

During the mass spectrometric measurements the species AlN⁺ was detected several times. Its IE value of 15.8 \pm 0.6 eV, obtained from the linear extrapolation method, was relatively large due to fragmentation of the parent molecule AlNC.²⁰ In order to reduce the fragmentation contribution to AlN⁺ we measured the ion intensity of AlN with 13 V electrons as 7.8E-14 A, whereas the ion intensity of AlNC⁺, involved in one of the reactions used, measured at 30 eV (see

Table II, Series 1) was corrected to 13 eV as 5.7E-9 A using its ionization efficiency curve (IEC). At these experimental conditions the ion current of AlN is 2% of that one of Al₂N, 4.0E-12 A, and there could still be a fragmentation contribution caused from Al₂N. The following reactions have been evaluated: AlN(g) + C(graph.) = AlNC(g) and AlN(s) = AlN(g), according to the third-law method.

The GEF₀ and HCF₀ for AlN(g), needed in the evaluation of the reactions enthalpies, were calculated using the vibrational frequency of 747 cm⁻¹ and the bond distance of 1.786 Å, by Simmons and McDonald.⁷ Two excited electronic states and respective energies, in cm⁻¹, based on theoretical computations of Langhoff *et al.*¹⁰ were also taken into account: ³Σ⁻(378), ¹Σ⁺(4689). Table III lists the thermal functions of AlN.

The relative ionization cross sections, σ , in 10⁻¹⁶ cm², used at 13 eV were: Al, 8.47;³⁹ C, 0.267;⁴⁰ N, 0.092;⁴⁰ AlN, 6.42 as 0.75 \times ($\sigma_{\text{Al}} + \sigma_{\text{N}}$); and AlNC, 6.62 as 0.75 \times ($\sigma_{\text{Al}} + \sigma_{\text{N}} + \sigma_{\text{C}}$). The multiplier gain of AlN was estimated to be 10⁵ times 4.

The $\Delta_f H_0^\circ$ values, in kJ mol⁻¹, obtained were -141 and 754. From these and auxiliary literature data, the $\Delta_a H_0^\circ$ values, in kJ mol⁻¹, were derived as 377 and 359, the average value being $\Delta_a H_0^\circ(\text{AlN}, g) = (368 \pm 15)$ kJ mol⁻¹, or (3.80 \pm 0.16) eV. The corresponding theoretical values, in eV, are: 2.60 (B3LYP);⁸ 2.35 (MRCI);¹⁰ 3.54 (LSDA);¹¹ 2.78 (GGA);¹¹ 2.45 (CCSD(T));¹⁴ 2.44 (G2);¹⁵ and 2.59 (CBS-RAD).¹⁵ Chase *et al.*³⁶ estimated the $D_0^\circ(\text{AlN}, g)$ of (2.85 \pm 0.37) eV from spectroscopic evidence.

Our upper limit for the dissociation energy of AlN is markedly higher than these computed values, except for the LSDA value; it is, however, consistent with the experimental values, in eV, for the isoelectronic molecules BP (3.56 \pm 0.18),⁴³ MgO (3.47 \pm 0.26),³⁶ and BeS (3.63 \pm 0.21),³⁵ which are somewhat lower, and with that of SiC (4.48 \pm 0.22),⁴⁴ which is higher. Isoelectronic molecules are known to often have similar atomization energies. Our upper limit value definitely shows that the dissociation energy of AlN is smaller than that of SiC.

The Gaussian-2 (G2) value of Swihart band Catoire¹⁵ is considerably lower than our upper limit. The G2 method is commonly thought reliable to within \pm 8.4 kJ mol⁻¹,⁴⁵ but as Schulz *et al.*⁴⁶ showed, results from G2 method can be very inaccurate in determining the thermodynamic properties of small polar metal-containing species. The reasons of this inaccuracy can be several, such as relaxation of inner-valence orbitals,⁴⁶ additivity approximation of energies,⁴⁶ and spin contamination.⁴⁷

A decision whether, e.g., the G2 theoretical value is accurate or whether our estimated upper limit is closer to the actual dissociation energy should come from new advanced spectroscopic measurements.

Al₂N₂

From a sensitive test to detect the Al₂N₂ molecule, we could estimate an upper limit of its ion intensity as 4E-14 A at 1940 K with 30 V electrons. For the determination of an upper limit of the enthalpy of atomization, $\Delta_a H_0^\circ$, of Al₂N₂

we evaluated the pressure independent reaction $2\text{Al}_2\text{N}(g) + \text{Au}(g) = \text{Al}_2\text{N}_2(g) + \text{AlAu}(g) + \text{Al}(g)$, according to the third-law method.

The ion currents for the other reactants are given in Table II, Series 1. The thermal functions, GEF_0 and HCF_0 , for $\text{Al}_2\text{N}_2(g)$ were computed using the molecular constants derived by Boo and Liu¹³ at the CCSD(T)/cc-pVTZ//MP2/cc-pVDZ level of theory for the ${}^1A_g(D_{2h})$ rhombic structure. The linear AlNNAI structure with a ${}^3\Sigma_g^-$ state by 50 kJ mol^{-1} above the ground state was also considered. Table III lists the resulting thermal functions of Al_2N_2 . The relative ionization cross section, in 10^{-16} cm^2 , at 30 eV for Al_2N_2 was taken as 15.9, as $0.75x(2\sigma_{\text{Al}} + 2\sigma_{\text{N}})$.

The resulting $\Delta_r H_0^\circ$ was equal to -127 kJ mol^{-1} and, using the auxiliary literature data listed above, the upper value of $\Delta_u H_0^\circ(\text{Al}_2\text{N}_2, g)$ was derived as 1402 kJ mol^{-1} . This value compares with the value of 1117 kJ mol^{-1} calculated by Andrews *et al.*,⁸ and with the value of 1070 kJ mol^{-1} by Boo and Liu.¹³ The latter value has been obtained using the total energy of Al_2N_2 , including the ZPE, given by Boo and Liu and the relative total energies of Al and N, computed by us at the same level of theory, CCSD(T)/cc-pVTZ, used by Boo and Liu. The computed total energies, in hartree, are -241.9312036 for Al, and -54.5147074 for N.

The failure of detecting gaseous Al_2N_2 is consistent with the available theoretical predictions, but it is difficult to reconcile with the report of Gordienko *et al.*⁴⁸ that gallium nitride vaporized from an open crucible mainly in the form of dimers, Ga_2N_2 . Specifically, it is concluded that the observation of Gordienko *et al.*⁴⁸ cannot correspond to equilibrium conditions among the vapor species above solid GaN.

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